

VISIBLE LIGHT pH CHANGE FOR ACTIVATING POLYMERS AND OTHER pH  
DEPENDENT REACTANTS

BACKGROUND OF THE INVENTION

*Ans*  
This invention relates generally to a method and apparatus for producing a pH change in a solution. More specifically, the invention relates to producing a pH change in a solution by irradiating the solution with visible light. With greater specificity, but without limitation thereto, the invention relates to using light to alter the pH of a solution to thereby cause an expansion and/or contraction of a pH dependent polymer immersed in the solution.

There exist a number of natural and synthetic fibers and gels that are expandable and contractible in volume when activated by an environmental change, such as exposure to a change in solvent composition, temperature, pH, electric field or photo irradiation, for example. As a commercially exploitable technology, the fibers and gels have applications in many fields, such as, for example, use in sensors, switches, motors, pumps, non-metallic operations and use in the medical and robotic fields where it is envisioned that these materials will be able to carry out the function of human muscle tissue.

The work of W. Kuhn and B. Hargitay as described in "Muskelähnliche Arbeitsleistung Kunstlicher Hochpolymerer Stoffe", Z. Elektrochemie 1951, 55(6), 490-502, incorporated by reference herein is one example of a synthesized polymer material capable of expansion and/or contraction. When the Kuhn and Hargitay polyacrylamide fiber, known as polyacrylic acid-polyvinyl alcohol (PAA-PVA), is placed within a solution of appropriately increasing pH, a 10% increase in fiber length is claimed to be observed.

Similarly, the work of T. Tanaka, D. Fillmore, S-T. Sun, I. Nishio, G. Swislow, and A.

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1 Shah described in the article "Phase Transitions in Ionic Gels" Phys. Rev. Lett. 1980, 45(20),  
2 1636-1639, incorporated by reference herein discloses an observed 400% volume collapse for a  
3 polyacrylamide gel disposed in a 50% acetone-water solvent mixture in which the pH of the  
4 solvent is lowered at constant temperature and solvent composition.


5 The work of Kuhn and Hargitay as well as Tanaka and Fillmore et al use a typical  
6 approach to changing the pH of a solution. In this approach, the pH is changed by manually  
7 dripping an acid or base into the solution. This technique, known as the "acid drip" method,  
8 relies upon the rate of the diffusion of hydrogen ions to a polymer site and is considered  
9 undesirably slow for certain polymer applications, such as use in synthetic muscles.

10 Besides the pH activation method of Kuhn and Hargitay and Tanaka and Fillmore et al,  
11 there exist electrical polymer activation schemes in which p-electron conjugated conducting  
12 polymers and electronically doped non-conducting polymers are electrically activated (expanded  
13 and contracted). An example of this activation method has been characterized by Shahinpoor et  
14 al as described in the article of D.J. Segalman, W.R. Witkowski, D.B. Adolf, and M. Shahinpoor  
15 titled: "Theory and Application of Electrically Controlled Polymeric Gels", Smart Materials and  
16 Structures, Vol. 1 (no. 1), M.V. Gandhi and B.S. Thompson (Eds.), London: Chapman and Hall,  
17 1992, 95-100. Like the pH activation method described above, the Shahinpoor et al method  
18 depends on the slow diffusion of ions to the active site of a polymer and therefore is also  
19 considered too slow for certain polymer applications such as use in synthetic muscles.

20 In addition to the activation approaches described above, there exist optical activation  
21 methods for causing volume changes in polymer fibers and gels. Noteworthy of these is the

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1 work of M. Irie and D. Kunwatchakun described in "Photoresponsive Polymers. 8. Reversible  
2 Photostimulated Dilation of Polyacrylamide Gels Having Triphenylmethane Leuco Derivatives",  
3 Macromolecules 1986, 19(10), 2476-2480. The Irie-Kunwatchakun studies were among the  
4 earliest on photoinduced volume changes in polymer gels. Photosensitive molecules, such as  
5 leucocyanide and leucohydroxide, were incorporated directly into a polymer's network.  
6 Irradiation with UV light produced a 2.2-fold reversible dimension change, but no significant  
7 volume change (phase transition) took place in the polymer studied, as the UV light-induced pH  
8 change was far from the pH null point of the polymer gel. Thus the magnitude of the dimension  
9 change was not optimized for certain applications such as robotics.

10 In the work of researchers Mamada and Tanaka as described in A. Mamada, T. Tanaka, D.  
11 Kungwatchakun, and M. Irie in "Photoinduced Phase Transition of Gels", Macromolecules 1990,   
12 23, 1517-1519 and as described in A. Mamada, T. Tanaka, D. Kungwatchakun, and M. Irie in  
13 United States Patent Number 5,242,491 titled: "Photo-Induced Reversible, Discontinuous  
14 Volume Changes in Gels" and issued Sep 7, 1993, photoinduced phase transitions in gels were  
15 observed. The copolymer used was that of Irie-Kunwatchakun described above. At a given  
16 temperature, the polymer gel discontinuously swelled in response to UV irradiation and shrank  
17 when the UV light was removed. It is hypothesized that this swelling is due to dissociation into  
18 ion pairs, thereby increasing internal osmotic pressure within the gel. The shrinking process of  
19 this method is governed by ion diffusion and recombination, making the speed of the reverse  
20 process impossible to control, thereby hindering its usefulness in many polymer actuator  
21 applications.

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1 In either of the UV studies described above, the UV radiation can cause undesired  
2 ionization, photolysis and molecular ligation of a utilized polymer.

3 Finally, in the work of A. Suzuki and T. Tanaka described in the article "Phase Transition  
4 in Polymer Gels Induced by Visible Light", Lett. Nature 1990, 346, 345-347, visible light was ←  
5 used to irradiate a gel containing a light-sensitive chromophore located in the backbone of an  
6 expandable and contractible copolymer. The chromophore absorbed the light and the light  
7 energy was then dissipated locally as heat by radiationless transitions, the result of which  
8 increased the "local" temperature of the polymer. Unlike the UV studies, the polymer expansion  
9 is a rapid process and is due to the direct heating of the polymer network by light. Yet the  
10 process of returning the polymer to its original size requires cooling, which becomes increasingly  
11 difficult as the temperature of the surrounding solution approaches the temperature of the  
12 polymer. This reverse process is too slow for many polymer uses such as in synthetic muscles.

13 Because many reactions are based on either acid or base catalyzations, including those of  
14 the polymers described above, researchers have investigated various approaches to promoting  
15 rapid pH changes. Such has been the case of Anthony Campillo et al as described in the article  
16 by A.J. Campillo, J.H. Clark, R.C. Hyer, S.L. Shapiro, K.R. Winn, and P.K. Woodbridge titled:  
17 "The Laser pH Jump", Proc. Intl. Conf. Lasers '78, Orlando, FL, Dec 11-15, 1978, Chem. Phys.  
18 Lett. 1979, 67(2), 218-222; the article by A.J. Campillo, J.H. Clark, S.L. Shapiro, K.R. Winn,  
19 and P.K. Woodbridge, titled: "Excited-State Protonation Kinetics of Coumarin 102", Chem.  
20 Phys. Lett. 1979, 67(2), 218-222; the article by J.H. Clark, S.L. Shapiro, A.J. Campillo, K.R.  
21 Winn, titled: "Picosecond Studies of Excited-State Protonation and Deprotonation Kinetics. The

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1 Laser pH Jump", J. Am. Chem. Soc. 1979, 101(3), 746-748; and United States Patent Number  
2 4,287,035 issued to John H. Clark, Anthony J. Campillo, Stanley L. Shapiro, and Kenneth R.  
3 Winn on Sep. 1, 1981.

4 The work of Campillo et al relies on excited-state proton transfer reactions to change the  
5  $[H^+]$  of a solution by several orders of magnitude. Campillo et al used a picosecond spectroscopy  
6 tool to directly measure excited-state deprotonation-protonation reaction rate constants. To  
7 promote a pH change, a UV laser with a pulse width of 20 picoseconds was used to excite  
8 2-naphthol-6-sulfonate to a higher ( $S_1$ ) electronic state. From the measured rate constants,  
9 Campillo et al determined that the excited-state  $pK_a$  value was 1.9, as opposed to the  
10 ground-state value of 9.1. This 7-unit change in  $pK_a$  corresponds to a 7-order of magnitude  
11 increase in the acid dissociation constant,  $K_a$ . Campillo's findings are consistent with earlier  
12 studies which show that excited-state  $K_a$  values can differ from ground-state values by many  
13 orders of magnitude, see the disclosure of J.F. Ireland and P.A.H. Wyatt titled: "Acid-Base  
14 Properties of Electronically Excited States of Organic Molecules", Adv. Phys. Org. Chem. 1976,  
15 12, 131-221.

16 Campillo et al claim that a major use of their technique is initiation of acid-base catalyzed  
17 ground-state reactions. For example, the reactants A and B are present in solution at pH 7. The  
18 ground state reaction,  $A + B \rightarrow C$ , occurs only at pH 4. By exciting the Campillo et al "jump  
19 molecule", 2-naphthol-6-sulfonate, a subnanosecond jump from pH 7 to pH 4 can be achieved,  
20 thereby enabling the desired ground-state reaction. Referring to FIG. 1, a schematic state energy  
21 level diagram illustrates the path by which the "jump molecule" 2-naphthol-6-sulfonate travels to

1 produce the pH change described. The 2-naphthol-6-sulfonate is irradiated with UV light and is  
2 excited from ground state  $S_0$  to first excited singlet state  $S_1$ . Radiative decay (florescence) then  
3 occurs bringing the molecule back to its ground state.

4 A major drawback of the Campillo technique is the extremely short duration of the  
5 accompanying pH change, typically 10 nanoseconds. While Campillo proposes that the excited  
6 state duration, and hence pH change, could be prolonged through use of repetitious irradiation,  
7 such an irradiation would require a bombardment of photons on the order of a million times a  
8 second. An additional shortcoming of the Campillo technique, when utilized with expandable  
9 and contractible polymers such as those described above, is that the utilized UV radiation  
10 promotes undesirable polymer ionization, photolysis and other molecular ligation. Additionally,  
11 the extremely narrow illumination path (0.1 mm or 5D-6 cubic centimeters) provided by the  
12 utilized 266 nanometer laser is considered insufficient to effectively illuminate an  
13 expandable/contractible polymer to undergo an appreciable change in volume.

#### 14 SUMMARY OF THE INVENTION

15 The invention provides a method and apparatus of rapidly changing the pH of a solution  
16 by way of a pH jump molecule that is activated by visible light. An application of the present  
17 invention is the ground-state reaction of changing the volume of an expandable and contractible  
18 polymer for simulated muscle applications as well as for other applications.

19 To permit these applications, it is desirable (1) to use a source of excitation energy that is  
20 not harmful to a utilized polymer; (2) to produce an *in-situ* pH change in which hydrogen ions  
21 become rapidly present at a polymer site; (3) to sustain the resultant pH change long enough and

1 in a volume large enough for desired ground-state reactions to occur, for example, the fully  
2 reversible expansion and contraction of a polymer; and (4) to provide a mechanism for efficient  
3 dissipation of heat produced as a result of the source of excitation energy.

4 Candidate pH "jump molecules" considered suitable for providing sufficient polymer  
5 actuation (activation) should possess the following characteristics:

6 (1) the jump molecules should have long lifetimes at room temperature, e.g 10  
7 milliseconds;

8 (2) the jump molecule acidity constants should be grossly different in ground and triplet  
9 states, e. g., 7 orders of magnitude;

10 (3) the resultant pH change should go through the midpoint (pH null point) of the  
11 utilized polymer; and

12 (4) either the non-protonated or the protonated form of the jump molecule should absorb  
13 in the visible region of the spectrum.

14 In accordance with the present invention, an apparatus and method incorporating these  
15 desirable features are disclosed. The invention includes a pH jump molecule that permits visible  
16 light excitation to provide a long lasting pH change to a pH dependent polymer or other pH  
17 driven reactant. The attendant pH change occurs rapidly (in nanoseconds) and will last for the  
18 excited state lifetime of the jump molecule. Further irradiation by either a continuous wave or  
19 appropriately pulsed laser can sustain the pH change indefinitely. Heat resulting from the light  
20 activation is efficiently discharged by radiative decay through room temperature  
21 phosphorescence lifetimes existing on the order of milliseconds. Thus an expandable and

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1 contractible polymer can be made to respond rapidly to a change in pH while the radiant heat-  
2 release mechanism of the invention allows the polymer to return to its initial configuration in a  
3 millisecond time frame, suitable for a variety of useful applications, including robotics.

4 Accordingly, it is an object of this invention to provide a method and apparatus for  
5 producing a rapid pH change in a solution.

6 A further object of this invention is to produce a rapid pH change in a solution that is  
7 useful in causing the expansion and/or contraction of a polymer.

8 Another object of this invention is to produce a rapid pH change in a solution that lasts  
9 long enough and is prevalent enough to be useful in causing the expansion and/or contraction of  
10 a polymer.

11 Still another object of this invention is to produce a rapid pH change in a solution that is  
12 useful in causing the expansion and/or contraction of a polymer while minimizing damage to the  
13 polymer.

14 Still yet another object of this invention is to produce a rapid pH change in a solution by  
15 irradiating the solution with visible light.

16 Yet another object of this invention is to produce a pH change in a solution by irradiating  
17 the solution with visible light in which any heat produced by the light is rapidly dissipated.

18 Other objects, advantages and new features of the invention will become apparent from  
19 the following detailed description of the invention when considered in <sup>conjunction</sup> ~~conduction~~ with the  
20 accompanying drawings.

21 BRIEF DESCRIPTION OF THE DRAWINGS



FIG. 1 is a schematic state energy level diagram.

FIG. 2 is illustrates the pH expansion and contraction dependence of an exemplary polymer, in this case an acrylamide gel.

FIG. 3 describes  $\Delta pK$  values for various families of molecules

FIG. 4 illustrates the light absorbance of anthracene versus wavelength.

FIG. 5 illustrates the pH expansion and contraction dependence of another exemplary polymer, in this case a polyacrylic acid- polyvinylalcohol (PAA-PVA) fiber.

#### DESCRIPTION OF THE PREFERRED EMBODIMENT

In the expandable and contractible polymer world, a term of art has evolved that describes the large and easily perceptible change in volume that occurs when such a polymer, whether it be a gel or a fiber, is exposed to a particular change in the pH of a solution in which the polymer is immersed. This term of art is known as a "phase transition", and describes the physical phenomenon that takes place when the polymer is exposed to a narrow change in pH that passes through what is know as the pH null point of the polymer.

Referring to FIG. 2, there is shown a graphical depiction of such a phase transition. This illustration, taken from the 1980 Physical Review Letter of T. Tanaka and D. Fillmore et al referred to above, shows the response of a polymer network of an acrylamide gel that has been hydrologized in a 4% (volume) N,N,N, N-tetramethylethylenediamine (TEMED) solution. The quantity  $\phi/\phi^*$  represents the swelling ratio which is the ratio of the final polymer network concentration to the initial polymer network concentration. The smooth curve is for gels immersed in water. The discontinuous curve is for gels in a 50% acetone-water mixture. In

1 either case, as pH is increased, the gel swells; as the pH is decreased, the gel shrinks.

2 For the acetone-water mixture shown in FIG. 2, the sharp s-shaped curve is characteristic  
3 of a phase transition. This behavior is referred to as a phase transition because an enormous

4 amount of polymer swelling-shrinking occurs within a very narrow range of pH values.

5 Capitalizing on this phenomenon, the greatest leverage for polymer activation can be achieved by

6 finding a polymer-polymer activation system that has a  $pK_a$  at the midpoint of the pH curve (or

7 what is otherwise referred to as the null point of the polymer). The closer that the ground state

8  $pK_a$  of a candidate "jump molecule" is to the null point of a polymer, the greater will be the

9 variability of polymer volume for a given quantity of excitation energy. By using such a jump

10 molecule, a small change in pH to either side of the midpoint will expand or contract the polymer

11 by the largest amount possible, optimizing polymer dimensional change for use in robotics or

12 other applications.

13 The term  $pK$  is a shorthand indicating the strength of an acid ( $pK_a$ ) and is defined as the  
14  $-\log_{10} K$  in which  $K$  is the characteristic equilibrium constant  $K$ , represented by:

15 
$$K = [H^+][B]/[BH^+]$$

16 where  $[H^+]$  is the hydrogen ion concentration and  $[B]$  is the concentration of the conjugate base.

17 When the amount of one of these constituents is varied, the others will adjust to keep  $K$   
18 constant.

19 During the course of scientific research, the inventor constructed kinetic equations for the  
20 3-level system of FIG. 1. Referring again to FIG. 1, an ideal "jump molecule" will be excited  
21 from ground state energy level ( $S_0$ ) to first excited singlet state energy level ( $S_1$ ), and return to

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1 the ground state via triplet state energy level ( $T_1$ ). The radiationless transition and radiative  
2 decay via phosphorescence will function as a "sink" for the molecules and because of their  
3 combined long lifetime, a prolonged molecule excited state will exist. The pH change produced  
4 by this excitation will last for the life of this excited state.

5 Repeated runs with many different candidate jump molecules predicted the requirements  
6 necessary to sustain a desired pH change:

7 (1) jump molecules should have long excited state lifetimes at room temperature, e.g.,

8 10 milliseconds;

9 (2) jump molecule acidity constants must be grossly different in the triplet and ground states,

10 e.g., 7 orders of magnitude;

11 (3) the resultant pH change should go through the midpoint (pH null point) of a utilized

12 polymer; and

13 (4) either the non-protonated or the protonated form of the jump molecule should absorb in the  
14 visible region of the spectrum. ←

15 A great many molecules with functional groups were eliminated based upon being  
16 disqualified by the above requirements.

17 For example, the phenones are considered undesirable because the lifetimes of the  
18 protonated and non-protonated forms are very different, providing a rapid excited state  
19 deactivation channel. An example of this is benzophenone, having an unprotonated lifetime of  
20 100 milliseconds and a protonated lifetime of 62 nanoseconds.

21 In addition, a great many functional groups were eliminated based upon small  $\Delta pK$

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values,  $\Delta pK$  in this instance being the difference between first triplet state  $pK$  value minus the ground state  $pK$  value ( $pK(T_1) - pK(S_0)$ ), as can be seen in FIG. 3.

In Table 1, characteristics of the carbon acids are described. The carbon acids shown exhibited long excited-state lifetimes  $\tau_p$  ( $p$  for phosphorescence), large  $\Delta pK$  values, and have  $\Delta pK$  values that pass through a desired polymer null point, however the excitation wavelength  $\lambda_{00}$  necessary to initiate a pH change falls within the ultraviolet. In this table, "obs" means "observed" and "c" means "calculated".

TABLE 1

	$\tau_p$ (msec)	$pK(S_0)$	$pK(S_1)$	$pK(T_1)$	$\lambda_{00}$ (nm)
fluorene	0.35	23.04	-5.96(c)	7.54(c)	300
9-phenylfluorene	obs	18.6	-10.7(c)	4.2(c)	305
9-cyanfluorene	obs	11.4	-12.4(c)	5.0(c)	300

Through the process of elimination, several families of molecules satisfied the pH jump molecule requirements stated above. One of these are the polynuclear aromatic hydrocarbons (PAC's) which are bases.

Of these, the PAC, anthracene, fits well with certain well established polymers. Referring to FIG. 4, the protonated form of this molecule is confirmed. In FIG. 4, an absorbance versus wavelength profile shows the zero-time spectrum for protonated anthracene. The peak at 424 nm is the only peak within the visible region of the spectrum which decreases with time, and is the signature of anthracene's protonated form. It is this peak that is used to activate the

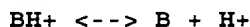
anthracene polymeric actuator with visible light.

Referring to FIG. 5, the contractile-expansion characteristics of the Kuhn-Hargitay polyacrylic-acid - polyvinylalcohol (PAA - PVA) polymer are shown. The Kuhn-Hargitay polymer fiber undergoes a phase transition between pH levels of 5 and 5.5, having a pH null point of approximately 5.3, as shown by the "Lange des Fadens" or "Length of Fiber" solid line.

Referring now to Table 2, specifications for utilizing protonated anthracene in coordination with the polymer described by Kuhn-Hargitay referred to above are shown.

TABLE 2

pH change and species concentrations  
BH<sup>+</sup> only absorbing, pH =5.0, 413.1 nm



## Anthracene Jump Molecule

		Ground	Singlet	Triplet	Lamda = 413.1 nm	
		-----	-----	-----	Log(eps) epsilon	
	pK's	3.8	13.6	10.3		
	Lifetimes nS (mS)		10.0	(10.0)		
	B	9.7D-4	1.5D-21	3.0D-11	0.04	1
	BH <sup>+</sup>	6.4D-4	2.0D-11	2.0D-4	4.38	23988

Initial Concentrations:

Final pH:

pH 5.0

5.48

[H<sup>+</sup>] 1.0D-5

3.3D-6

[B] 9.8D-4

[BH<sup>+</sup>] 2.0D-4

Watts 4.2

Photons/sec 9.3D+18

Total B 1.0D-3

V cm3 1.0

P/cm3-sec 9.3D+18

413.1 nm = Center Kr<sup>+</sup> line: 406.7, 413.1, 415.4

By utilizing visible light, the protonated form BH<sup>+</sup> of anthracene is disassociated into its base (B)

1 and hydrogen ion ( $H^+$ ) constituents to prompt a pH change from 5 to 5.48. As can be seen, the  
2  $\Delta pK$  ( $pK(T_1) - pK(S_0)$ ) of anthracene is 10.3 - 3.8, permitting such a large scale pH change.  
3 The calculation in Table 2 is based on a  $pK(S_0)$  value for anthracene found in Mackor.. E.L.,  
4 Hofstra, A., and Van Der Waals, J. H., 1958, in an article entitled "The Basicity of Aromatic  
5 Hydrocarbons", Trans. Faraday Soc., vol. 54, 66.

6 For use with the referenced Kuhn-Hargitay polymer, the desired protonated form of  
7 anthracene is derived by dissolving enough anthracene in cyclohexane, as described in Table 2,  
8 so that the resulting concentration of non-protonated anthracene is  $9.8 \times 10^{-4}$  moles/liter when the  
9 pH is adjusted to 5.0 by the addition of sulfuric acid ( $H_2SO_4$ ). The mixture is then vigorously  
10 shaken in a separatory funnel, causing the anthracene to diffuse from the cyclohexane to the  
11 sulfuric acid to form a solution of protonated anthracene.

12 For the polymer-anthracene combination described, a BeamLok 2080 krypton ion laser  
13 was used to irradiate the polymer system at 413.1 nanometers and 4.2 watts. The one cubic  
14 centimeter irradiation volume is large enough to house a polymer of macroscopic dimensions as  
15 the jump molecule provides a pH change from 5.0 to 5.48. Because of the 10 millisecond  
16 prolonged excited state of the anthracene jump molecules, the continuous wave laser will permit  
17 constant pH elevation until the irradiation is cut-off, at which time the excited-state jump  
18 molecules will decay to the ground state and reassociate, causing a return to the original pH in a  
19 few milliseconds. Importantly, the heat created by the molecules absorbing the irradiated light is  
20 released as light of a longer wavelength. Full polymer reversibility, which is not hindered by the  
21 slow dissipation of heat, is therefore made possible for use in many polymer applications,

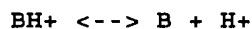
including robotics.

Besides use of a continuous wave irradiation source, a pulsed laser having a repetition rate of 100 times a second at 42 millijoules will also suffice. This repetition rate will prompt a pulse every 10 milliseconds, permitting continuous pH elevation.

Referring now to Table 3, specifications for utilizing protonated anthracene in coordination with the polymer described by Tanaka-Fillmore et al referred to above are shown. The protonated form  $BH^+$  of anthracene coordinates well with the Tanaka polymer in which the null point of this polymer (3.8 pH) corresponds with the ground state pKa value of the anthracene.

TABLE 3

pH change and species concentrations  
 $BH^+$  only absorbing, pH = 3.7, 413.1 nm



## Anthracene Jump Molecule

		Ground	Singlet	Triplet	Lamda = 413.1 nm	
		-----	-----	-----	Log(eps) epsilon	
	pK's	3.8	13.6	10.3		
	Lifetimes nS (mS)		10.0	(10.0)		
	B	6.4D-4	3.2D-21	6.9D-11	0.04	1
	BH+	1.8D-4	2.3D-11	1.8D-4	4.38	23988
Initial Concentrations:		Final pH:				
	pH	3.7	3.9			
	[H+]	2.0D-4	1.3D-4			
	[B]	7.1D-4				
	[BH+]	2.9D-4	Watts	6.3	Photons/sec	1.4D+19
	Total B	1.0D-3	V cm3	1.0	P/cm3-sec	1.4D+19

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413.1 nm = Center Kr+ line: 406.7, 413.1, 415.4

For use with the Tanaka-Fillmore polymer, the desired protonated form of anthracene is derived by dissolving enough anthracene in cyclohexane, as described in Table 3, so that the resulting concentration of non-protonated anthracene is  $7.1 \times 10^{-4}$  moles/liter when the pH is adjusted to 3.7 by the addition of sulfuric acid ( $\text{CH}_3\text{SO}_3\text{H}$ ). The mixture is then vigorously shaken in a separatory funnel, causing the anthracene to diffuse from the cyclohexane to the sulfuric acid to form a solution of protonated anthracene.

For the polymer-anthracene combination described, a BeamLok 2080 krypton ion laser may be used to irradiate the polymer system at 413.1 nanometers and 6.3 watts. The one cubic centimeter irradiation volume is large enough to house a polymer of macroscopic dimensions as the jump molecule provides a pH change from 3.7 to 3.9. The 10 millisecond prolonged excited state, permits the continuous wave laser to maintain a constant elevated pH level. Once the irradiation is cut-off, the excited-state jump molecules will decay to the ground state and reassociate, causing a return to the original pH in a few milliseconds. As before stated, the heat created by the jump molecules absorbing light will be efficiently discharged as light of a longer wavelength.

Obviously, many modifications and variations of the invention are possible in light of the above teachings. It is therefore to be understood that within the scope of the appended claims the invention may be practiced otherwise than as has been described.